

Anticipation Guide : Electrons and Chemical Bonding

Name _____

Date _____

Before reading pages 328-331: In the space to the left of each statement, place a check mark (✓) if you agree or think the statement is true.

During or after reading: Add new check marks or cross through those about which you have changed your mind. Keep in mind that this is not like the traditional “worksheet.” You may have to put on your thinking caps and “read between the lines.” Use the space under each statement to note the page(s), and paragraph(s) where you are finding information to support your thinking.

___ 1. When atoms join together to form new substances, the properties of the new substances are different from those of the substances that were combined.

___ 2. A theory is generally accepted as truth because it always proves true in experimental tests.

___ 3. To be able to tell how an atom will combine with another atom, you just need to know how many electrons each atom has.

___ 4. Valence electrons are the key to bonding since they are the ones in the outermost level of an atom, and they are most likely to come into contact with other atoms.

___ 5. If an atom has 8 valence electrons, it will not bond with other atoms.

___ 6. An atom with 6 outer-shell electrons might bond with two atoms that have only one outer-shell electrons, or it might bond with one atom having two outer-shell electrons.

___ 7. You can tell how an atom might combine with other atoms just by looking where it is located in the Periodic Table.

___ 8. An atom with five valence electrons could combine with either one, two, or three atoms of another substance.